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~~in Foods Calorimetry Lab~~ 22. Heat of Reaction of HCl V NaOH Calorimetry Experiment with different metals

~~CALORIMETRY Part 01~~

Specific Heat Capacity Experiment
Determining the enthalpy of solution of sodium hydroxide

Determination of Water Equivalent of a

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Calorimetry Lab and Calculations

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Student Exploration: Calorimetry Lab
Vocabulary: calorie, calorimeter, joule, specific heat

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B. Suppose a snack bar is burned in a calorimeter and heats 2,000 g water by 20 °C. How much heat energy was released? (Hint: Use the specific heat equation.) Give your answer in both joules and calories.

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~~Student Exploration Calorimetry Lab
(ANSWER KEY)~~

Student Exploration: Calorimetry Lab
Vocabulary: calorie, calorimeter, joule, specific heat capacity
Prior Knowledge Questions (Do these BEFORE using the Gizmo.)
1. The Latin word calor means "heat," and meter comes from

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the Greek word meaning "to measure."
What do you think a calorimeter does?
Measures heat 2. Where have you
heard the word calorie before? ? What
do you think a calorie

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$c_p(\text{H}_2\text{O}) = 4.184 \text{ J / g } ^\circ\text{C}$ $q = mc_p \Delta T$

1) A compound is burned in a bomb calorimeter that contains 3.00 L of water. If the combustion of 0.285 moles of this compound causes the

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temperature of the water to

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Substitution into equation 3 yields equation 4: (4) $q_{\text{water}} = m_{\text{water}} \cdot s_{\text{water}} \cdot \Delta T_{\text{water}}$. Finally, equations 1, 2 and 3 can be combined into the calorimetry equation: (5) $q_{\text{sys}} = -m_{\text{water}} \cdot s_{\text{water}} \cdot \Delta T_{\text{water}} = -m_{\text{water}} \cdot (4.18 \text{ J/g} \cdot ^\circ\text{C}) \cdot \Delta T_{\text{water}}$.

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~~Lab 4 - Calorimetry~~

Correct Answer: C. The final temperature of the lead-water system will be lower than the final temperature of the copper-water system. A blacksmith heats a 1,540 g iron horseshoe to a temperature of 1445°C

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before dropping it into 4,280 g of water at 23.1 °C.

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Soda Can Calorimeter Energy Content of Food ... This lab should be performed in a well-ventilated room.

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Wash hands thoroughly ...

(kilocalories) by dividing the answer above by 1000. $265 \text{ cal.} \div 1000 =$

0.265 Cal. 4. Determine how much of the food burned by subtracting the final mass of the cork/pin/food assembly from the initial mass.

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~~Soda Can Calorimeter - Flinn Scientific~~

6.03 Calorimetry Lab Report By;

Selina Pfuner CALCULATIONS P2

Unknown Metals $q[\text{water}] = m \times c \times$

ΔT $m = 24.5 \text{ g}$ $c = 4.18$ $\Delta T =$

$29.1 - 25.2 = 3.9$ $\Delta T = \text{final temp} - \text{initial}$

temp $24.5 \times 4.18 \times 3.9 = 399 \text{ J}$ multiply

all together equal $q[\text{water}]$ A B C

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-399J=25.605[29.1-100.5]

~~6.03 Calorimetry Lab Report by Selina Pfuner - Prezi~~

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not forget to include any additional details, which may be useful for readers. Composing a great science lab sample is essential if you wish to make your research and your report intriguing and beneficial to readers.

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Sample

Online Library Answer Sheet For Calorimetry Lab Worksheet Key Open the Calorimetry Lab Activity (Section 6.6, Figure 6.11) to complete this worksheet. 1. Determine the masses of the cup, the water, and the gold block. Object Cup Water Gold Mass

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(g) 4.0g 60.0g 38.69 2. Measure the initial temperature of the water.

Solved: Chapter 6 Calorimetry Worksheet Key Open

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EXPERIMENT Heat Of Neutralization
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Question: Name: LAB #4 -

CALORIMETRY Calorimetry Is The Practice Of Measuring The Heat Released By A Chemical Reaction By Observing The Effect Of That Heat On A Substance Of Known Specific Heat

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(usually Water). If A Known Mass Of Water Undergoes A Measurable Temperature Change As A Result Of A Chemical Reaction, The Amount Of Heat Released By The Reaction Can ...

~~Solved: Name: LAB #4~~

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~~CALORIMETRY Calorimetry Is The Prac ...~~

Question: DATA SHEET For
CALORIMETRY EXPERIMENTS A)
SPECIFIC HEAT Of A METAL METAL
UNKNOWN Number Trial 2 (50 MI
Water) 13.4502 23.9235 100.0 5.0250
55.020 Trial 1 (40 MI Water) Mass Of

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Empty Test Tube (tare) 13.4502 Mass
Of Tube + Metal 23.7502 Temperature
Of Boiling Water 100.0 Mass Of Empty
Calorimeter 5.0250 Mass Of
Calorimeter + Water 45.0250 Initial
Temperature ...

~~DATA SHEET For CALORIMETRY~~

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~~EXPERIMENTS A) SPECIFIC ...~~

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~~Virtual Lab~~

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Investigate how calorimetry can be used to find relative specific heat

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values when different substances are mixed with water. Modify initial mass and temperature values to see effects on the system. One or any combination of the substances can be mixed with water. A dynamic graph (temperature vs. time) shows temperatures of the individual

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substances after mixing.

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Calorimetry POGIL.notebook 9

December 06, 2012 Key Equaon (try to answer number 13 and if you are not sure, ask!) 13.To change the

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